

## Concentration and Solubility

### Understanding Main Ideas

Answer the following questions on the lines below.

1. What amounts do you compare when measuring concentration?

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2. How can you tell that a white powder is salt without tasting it?

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3. Which solution will have more gas dissolved in it, a solution under high pressure or one under low pressure?

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4. Explain the meaning of the expression “like dissolves like.”

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5. How does temperature affect the solubility of most solids?

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### Building Vocabulary

Match each term with its definition by writing the letter of the correct definition on the line beside the term in the left column.

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|-----------------------------------|--|
| _____ 6. dilute solution          | a. a measure of how much solute can dissolve in a solvent at a given temperature |
| _____ 7. concentrated solution    | b. a solution that has more dissolved solute than is predicted by its solubility |
| _____ 8. solubility               | c. a solution that has so much solute that no more dissolves                     |
| _____ 9. saturated solution       | d. a solution that has only a little solute                                      |
| _____ 10. unsaturated solution    | e. a solution in which more solute can be dissolved                              |
| _____ 11. supersaturated solution | f. a solution that has a lot of solute   |